metal-organic papers

Acta Crystallographica Section E Structure Reports Online

ISSN 1600-5368

Rafal Grubba,^a* Wieslaw Wojnowski,^a Katarzyna Baranowska,^a Elke Baum^b and Jerzy Pikies^a

^aFaculty of Chemistry, Gdańsk University of Technology, Narutowicza 11/12, Gdańsk PL-80952, Poland, and ^bInstitut für Anorganische Chemie der Universität Karlsruhe TH, Engesserstr. Geb. 30.45 D-76128 Karlsruhe, Germany

Correspondence e-mail: raf710@wp.pl

Key indicators

Single-crystal X-ray study T = 120 KMean $\sigma(\text{C}-\text{C}) = 0.005 \text{ Å}$ R factor = 0.040 wR factor = 0.097 Data-to-parameter ratio = 16.0

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

 $\ensuremath{\mathbb{C}}$ 2006 International Union of Crystallography All rights reserved

Bis(dithiobenzoato- $\kappa^2 S, S'$)bis(tetrahydrofuran- κO)magnesium(II)

The title compound, $[Mg(C_7H_5S_2)_2(C_4H_8O)_2]$, was obtained by the reaction of PhMgBr with an excess of CS₂. The magnesium(II) cation is six-coordinated by two dithiobenzoate ligands coordinating through their S atoms in a bidentate chelating fashion and two O atoms from two tetrahydrofuran ligands, in a distorted octahedral geometry.

Comment

In the course of our studies on silicon-sulfur chemistry, we have synthesized two magnesium derivatives of dithiobenzoic acid, *viz.* PhCS₂MgS₂CPh·2THF, (I), and PhCS₂MgBr·3THF, (II), by the reaction of PhMgBr with an excess of CS₂ (Ramadan *et al.*, 1983). The formation of (I) can be rationalized in terms of an equilibrium involving (II) or in terms of a Schlenk equilibrium (Schlenk *et al.*, 1929) involving PhMgBr. In the second case, the formation of (I) can be described as the reaction of magnesium diphenyl with CS₂. To the best of our knowledge, compound (I) is the first example of a magnesium salt with two dithiobenzoate ligands.



In complex (I) (Fig. 1), the magnesium(II) cation is sixcoordinated by four S atoms from two dithiobenzoate anions and two O atoms from two tetrahydrofuran molecules, in a distorted octahedral geometry. Deviations from the ideal geometry are imposed by the S-C-S angle of the chelating dithiobenzoate ligands (Table 1). Few metal complexes of Ni (Tang et al., 1984; Fackler et al., 1984), Fe (Coucouvanis et al., 1969), Cr (Tang et al., 1985) and Co (Lu et al., 1995), with two or more ligands containing the dithiocarboxy group and with geometry close to octahedral, have been reported. Among them, an Ni complex containing two α -naphthyldithiocarboxylate ligands and two molecules of pyridine (Tang et al. 1984) is similar to complex (I). The C-S distances are between those of a typical C–S single bond (1.819 Å) and a C=S double bond (1.599 Å) (Prince, 2004), indicating electron delocalization in the dithiocarboxy groups. No classical hydrogen bonds were observed in the crystal structure.

Received 19 July 2006 Accepted 2 August 2006



Figure 1

The molecular structure of (I), showing the atomic numbering scheme. Displacement ellipsoids are drawn at the 30% probability level. H atoms have been omitted for clarity.

Experimental

The work was carried out using standard vacuum-nitrogen line and Schlenk techniques. PhMgBr (100 ml, 0.06 mol) was added dropwise into CS_2 (26 ml, 0.36 mol) at room temperature and the resulting red solution was concentrated under vacuum. A small amount of oily residue was obtained which was diluted in toluene (2 ml). Pentane vapours were diffused into the toluene solution and the solution was left to stand at room temperature for two months, after which time red–orange crystals of (I) were deposited.

Crystal data

 $\begin{bmatrix} Mg(C_7H_5S_2)_2(C_4H_8O)_2 \end{bmatrix} \\ M_r = 474.98 \\ Monoclinic, P2_1/n \\ a = 8.3846 (9) Å \\ b = 14.3899 (9) Å \\ c = 20.0275 (15) Å \\ \beta = 90.287 (11)^{\circ} \\ V = 2416.4 (4) Å^3 \end{bmatrix}$

Data collection

Stoe IPDS diffractometer φ scans Absorption correction: numerical [X-RED32 (Stoe & Cie, 1997); crystal shape approximated by a polyhedron with 13 faces] $T_{\rm min} = 0.781, T_{\rm max} = 0.876$

Refinement

Refinement on F^2 $R[F^2 > 2\sigma(F^2)] = 0.040$ $wR(F^2) = 0.097$ S = 0.854201 reflections 262 parameters Z = 4 D_x = 1.306 Mg m⁻³ Mo K α radiation μ = 0.44 mm⁻¹ T = 120 (2) K Block, red–orange 0.55 × 0.45 × 0.3 mm

10590 measured reflections 4201 independent reflections 2373 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.039$ $\theta_{\text{max}} = 25.1^{\circ}$

H-atom parameters constrained $w = 1/[\sigma^2(F_o^2) + (0.0526P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{max} < 0.001$ $\Delta\rho_{max} = 0.31$ e Å⁻³ $\Delta\rho_{min} = -0.19$ e Å⁻³

ab	le	1				

Т

Selected geometric parameters (Å, $^\circ).$

S1-Mg1	2.5614 (11)	S4-Mg1	2.5753 (11)
S2-Mg1	2.6086 (13)	O1-Mg1	2.065 (2)
S3-Mg1	2.5831 (14)	O2-Mg1	2.074 (2)
S2-C1-S1	121.37 (16)	O2-Mg1-S3	92.78 (8)
S3-C8-S4	120.57 (19)	S1-Mg1-S3	97.95 (4)
O1-Mg1-O2	87.59 (10)	S4-Mg1-S3	69.30 (4)
O1-Mg1-S1	99.35 (7)	O1-Mg1-S2	90.61 (8)
O2-Mg1-S1	94.35 (6)	O2-Mg1-S2	163.13 (7)
O1-Mg1-S4	93.55 (7)	S1-Mg1-S2	69.39 (3)
O2-Mg1-S4	100.94 (6)	S4-Mg1-S2	95.91 (4)
S1-Mg1-S4	160.38 (5)	S3-Mg1-S2	93.90 (5)
O1-Mg1-S3	162.61 (7)		

H atoms were placed in idealized positions and allowed to ride on their parent atoms, with C-H = 0.95-0.99 Å and $U_{iso}(H) = 1.2U_{eq}(C)$.

Data collection: *IPDS Software* (Stoe & Cie, 1997); cell refinement: *IPDS Software*; data reduction: *X-RED32* (Stoe & Cie, 1997); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

Financial support to RG from the Polish Ministry of Research and Information (project No. 3T09A 120 28) is gratefully acknowledged.

References

Coucouvanis, D. & Lippard, S. J. (1969). J. Am. Chem. Soc. 91, 307-311.

- Fackler, J. P. Jr, Del Niera, R., Campana, C. & Trzcinska-Bancroft, B. (1984). J. Am. Chem. Soc. 106, 7883–7886.
- Farrugia, L. J. (1997). J. Appl. Cryst. 30, 565.
- Farrugia, L. J. (1999). J. Appl. Cryst. 32, 837-838.
- Lu, S.-F., Chen, H.-B. & Huang, X.-Y. (1995). *Jiegou Huaxue (Chin. J. Struct. Chem.)*, **14**, 261. (In Chinese.)
- Prince, E. (2004). Editor. International Tables for Crystallography, Vol. C, 3rd ed., Table. 9.5.1.1, pp. 806–807. Heidelberg: Springer.
- Ramadan, S. R., Srinivasan, P. S., Ramachadran, J. & Sastry, V. V. S. K. (1983). Synthesis, 6, 605–622.
- Schlenk, W. & Schlenk, W. Jr (1929). Chem. Ber. 62, 920-924.
- Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.
- Stoe & Cie (1997). *IPDS Software* and *X-RED*32. Stoe & Cie GmbH, Darmstadt, Germany.
- Tang, K., Gan, H., Pan, Z. & Jin, X. (1984). Beijing Dax. Xue. Zir. Kex. (Acta Sci. Nat. Univ. Pek.), 43, 3>. (In Chinese.)
- Tang, K., Yang, Q. & Li, Q. (1985). Beijing Dax. Xue. Zir. Kex. (Acta Sci. Nat. Univ. Pek.), 6, 12. (In Chinese.)